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ISOTOPES AND

AVERAGE Name

ATOMIC MASS

Elements come in a variety of isotopes, meaning they are made up of atoms with the same atomic number but different atomic masses These atoms differ in the number

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(metal + nonmetal),

covalent (nonmetal +

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Name NAMING

MOLECULAR

COMPOUNDS Name the

following covalent

compounds. 12.. 16.

17. 19. 20. So Cac IGde

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5. 35 g of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$

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Determine the number of grams in each of the quantities below. 1, 2.5 moles of NaCl 2. 0.50 moles of 3, 1.70 moles of KMnO_4 4, 0.25 moles of KCl e, ST 5. 3.2 moles of $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
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LE CHATELIER'S

PRINCIPLE Name _ Le

Chatelier's Principle

states that when a

system at equilibrium

is subjected to a stress,

the system will shift its

equilibrium point in

order to relieve the

stress. Stoichiometry

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ATOMIC STRUCTURE

AND THE PERIODIC

TABLE CHAPTER 6 & 7

WORKSHEET Read the

clue and give the

element symbol that

applies. CLUE 1. 2.

|He SAI

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A chemical equation is the symbol in Chemistry that represents chemical reaction with the help of chemical formulas. It contains the chemical substances that are involved in the reaction. It contains reactants and products. The reactants are the elements that react with one and another in a chemical reaction, while the products are

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Equilibrium Shift right
+ OH-(aq) Equilibrium

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10,6
kcal Amount NaOH(s)

Same Same 82

decreases same Same

remains the increases

same same same

Same Same same

same same

(Remember that pure

solids and liquids do

not affect equilibrium

values.) $[Na^+]$ same S_a

$[OH^-]$ same Same same

S_a me same same

same

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